

## Chapter 9 Stoichiometry Answers Section 2

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### Chapter 9 Stoichiometry Answers Section

CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$  a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of  $C_3H_4$ ? 2 mol O<sub>2</sub>:1 mol H<sub>2</sub>O c. What is the mole ratio of O<sub>2</sub> to H

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Chapter 9 - Stoichiometry 9-1 Introduction to Stoichiometry Composition Stoichiometry - deals with mass relationships of elements in compounds Reaction Stoichiometry - Involves mass relationships between reactants and products in a chemical reaction I. Reaction Stoichiometry Problems A.

### Chapter 9 - Stoichiometry

Modern Chemistry 77 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. \_\_\_\_ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H

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### CHAPTER 9 REVIEW

Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students...

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Stoichiometry. SECTION 2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many moles of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed? 2. Given the following equation:  $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

### CHAPTER 9 REVIEW

CHAPTER 9 REVIEW. Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_ The coefficients in a chemical equation represent the. (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of  $\text{N}_2$  are mixed with 120 mol of  $\text{H}_2$  according to the following equation:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

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